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Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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Gases Class SHORT ANSWER Answer the following questions in the space provided.

c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (b) divided by the mass of 1 mol. nRT (c) multiplied by 22.4 L. (d) divided by 22.4 L. For the expression $V =$ (a) increasing

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3 react according to the above equation, how many moles of NO will be consumed? 8. Propyne gas can be used as a fuel. The combustion reaction of propyne can be represented by the following equation: $C_3H_4(g) + 4O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. Write all the

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possible mole ratios in this system. 4
mol O 2:1 mol C 3H 4; 3 mol CO 2:1 mol
C 3H 4; 2 ...

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Periodic Law SECTION 1 SHORT ANSWER
Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass.

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Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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SECTION 3 continued 5. Explain the following properties of solids by describing what is occurring at the atomic level. a. Metallic solids conduct electricity well, but covalent network solids do not. Metals have many electrons that are not bound to any one

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atom; therefore they are able to move throughout the crystal.

10 States of Matter - Ms. Agostine's Chemistry Page

Modern Chemistry 84 . Name Section
Quiz, continued Class Date 5. In some
instances, the concentration of a
solution is expressed as molality instead

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of molarity because a. molality is easier to calculate. b. molarity applies only to solid-liquid solutions.

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CHAPTER 6 STUDY GUIDE Chemical
Bonding SECTION 3 IONIC BONDING AND
IONIC COMPOUNDS SHORT ANSWER

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Answer the following questions in the space provided. 1. _____ The notation for sodium chloride, NaCl, stands for one (a) formula unit.

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CHAPTER 3 REVIEW Atoms: The Building
Blocks of Matter SECTION 3 SHORT

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ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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