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 $2 \frac{1}{T} = \frac{\partial S}{\partial U}$ (7) $= \frac{1}{T} \frac{\partial S}{\partial q} = k [\ln(q+N)+1 \ln q^{-1}]$ (9) $= k \ln \frac{q+N}{q}$ (10) $= k \ln \frac{U+N}{U}$ (11) $T = k \ln \frac{U+N}{U}$ (12) We can invert this to obtain $U(T)$: $e = kT = U + N \frac{U}{U+N}$ (13) $U = N e^{-kT}$ (14) Finally, the heat capacity is obtained from $C = \frac{\partial U}{\partial T}$ (15) $= N e^{-kT} = kT^{-2} (16) = 2Ne^{-kT} = kT^{-2} e^{-kT} (17)$

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The thermal energy of a monatomic gas is $\frac{3}{2} NkT = \frac{3}{2} PV$ so the change in thermal energy due to the expansion is $\Delta U = \frac{3}{2} P_f V_f - \frac{3}{2} P_i V_i$ (9) $= \frac{3}{2} (1.013 \times 10^5 \text{ Pa}) (10 \text{ m}^3) - \frac{3}{2} (1.013 \times 10^5 \text{ Pa}) (3 \text{ m}^3)$ (10) $= 1215.6 \text{ J}$ (11) The heat added to the gas must be enough to do the work plus increase the thermal energy, so $Q = \Delta U + W = 1620.8 \text{ J}$ (12) Example 2.

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Thermodynamic identity (Schroeder 3.4) $dU = TdS - pdV$ (1.5)
Total differentials When we have a total differential, the things in front of the dS, dV , etc are partial derivatives. $T = \frac{\partial U}{\partial S} \quad -p = \frac{\partial U}{\partial V}$ (1.6) Also, you can integrate along a path using a total differential, and can do linear algebra using total differential equations, e.g. substituting one for another.

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